

Predicting Products Review Stations

Part I : Predicting Products

Directions: Predict the products of the reaction and balance the equation. For double replacement reactions, be sure to include the states of the products.

1. $\text{SrCl}_2 + (\text{NH}_4)_2\text{SO}_4 \rightarrow \text{SrSO}_4(\text{s}) + 2\text{NH}_4\text{Cl}(\text{aq})$
2. $\text{Ba}(\text{OH})_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaOH}(\text{aq})$

Directions: Predict the products of the reaction and balance the equation. For double replacement reactions, be sure to include the states of the products.

3. $\text{F}_2 + 2\text{NaBr} \rightarrow 2\text{NaF} + 2\text{Br}_2$
4. $\text{NaNO}_3 + \text{KOH} \rightarrow \text{NaOH}(\text{aq}) + \text{KNO}_3(\text{aq})$ NO RXN

Directions: Predict the products of the reaction and balance the equation. For double replacement reactions, be sure to include the states of the products.

5. $\text{CdCl}_2 + \text{Na}_2\text{CO}_3 \rightarrow 2\text{NaCl}(\text{aq}) + \text{CdCO}_3(\text{s})$
6. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$

Directions: Translate the equation, predict the products of the reaction and balance the equation. For double replacement reactions, be sure to include the states of the products.

7. Zinc chloride + Iodine \rightarrow

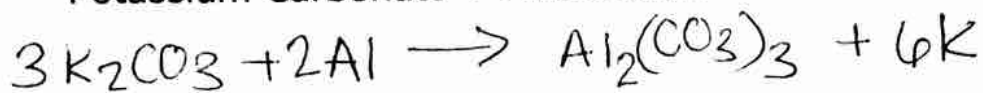


8. Calcium Acetate + Potassium Nitrate \rightarrow

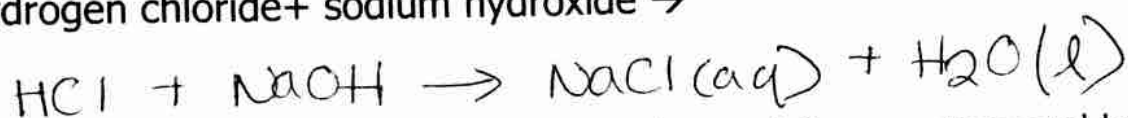


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9. Potassium Carbonate + Aluminum →

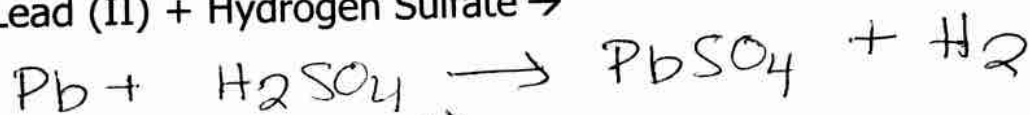


10. Hydrogen chloride + sodium hydroxide →

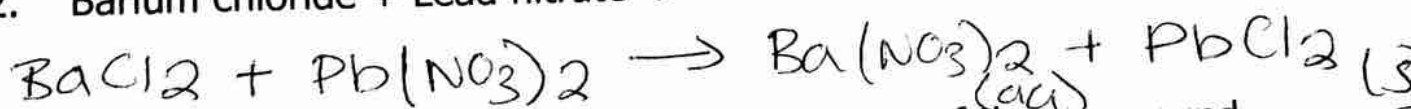


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11. Lead (II) + Hydrogen Sulfate →



12. Barium chloride + Lead^(II) nitrate →



Directions: Use your solubility rules to determine the state of the compound.

13. H_2S
gas

14. $MgCl_2$
aq

15. $FeCO_3$
solid

16. $(NH_4)_3PO_4$
aq