

**Bonding Review Stations**

Chemistry GT

**Station 1:** Write the name or the formula for the following **covalent compounds**:

1. Antimony tribromide SbBr<sub>3</sub>
2. Chlorine dioxide ClO<sub>2</sub>
3. Diphosphorus tetrahydride P<sub>2</sub>H<sub>4</sub>
4. SO<sub>2</sub> sulfur dioxide
5. N<sub>2</sub>H<sub>4</sub> dinitrogen tetra hydride
6. P<sub>2</sub>S<sub>3</sub> diphosphorous trisulfide

**Station 2:** Write the name or the formula for the following **covalent compounds**:

7. Iodine pentafluoride IF<sub>5</sub>
8. Dinitrogen trioxide N<sub>2</sub>O<sub>3</sub>
9. Phosphorus triiodide PI<sub>3</sub>
10. NF<sub>3</sub> nitrogen trifluoride
11. SF<sub>6</sub> sulfur hexafluoride
12. CO<sub>2</sub> carbon dioxide

**Station 3:** Write the name or the formula for the following **ionic compounds** using the **stock system**:

13. Potassium iodide KI
14. Barium chloride BaCl<sub>2</sub>
15. Lithium bromide LiBr
16. Na<sub>2</sub>CO<sub>3</sub> sodium carbonate
17. NaOH sodium hydroxide
18. MgBr<sub>2</sub> magnesium bromide